Chemistry Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Worksheeet – Formula Weight Calculations # 1

Directions: Using a periodic chart, find the correct Formula Weight.

PLEASE NOTE THE FOLLOWING:

1. When using the chart, you are allowed to round the atomic weights that are given on the chart. The limit of rounding is +/- .03 grams.
   1. Here are examples of what can be rounded:
      1. Oxygen to 16
      2. Hydrogen to 1
      3. Bismuth to 209
      4. Argon to 39.95 Notice that it can’t be rounded to 40.
   2. Here are some examples that can’t be rounded:
      1. Sulfur to 32
      2. Bromine to 80
      3. Boron to 11
      4. Magnesium to 24
2. Remember that the Coefficient is not a factor in determining formula weights.
   1. Examples:
      1. 2 NaCl has a formula weight of 58.45 g/n
      2. 5 H2O has a formula weight of 18 g/n
3. If you see a raised dot, treat it as an addition sign.
   1. Example:
      1. BaCl2 ● 2H­2O BaCl2 is 208.23 & 2 water molecules are 36 TOTAL = 244.23 g/n
4. A subscript after a ( ) counts for everything in the ( ).
   1. Example
      1. Cu(OH)2 The subscript 2 means that there are 2 Hydroxide Ions (OH-1) not just 2 hydrogens.

Solve the following problems on another sheet of paper.

1. 5 NaCl 2. GaBr3 3. AgNO3 4. Cr(SO4)3

5. RbCN 6. 3 CF4 7. CuSO4 ● 5 H2O 8. SO3

9. Zn(OH)2 10. 6 N2O3 11. Pb(C­2O4)2 12. 5 AsCl5